

CHEMISTRY

Learning objectives I can:	I can do this very well	I can do this quite well	I need to do more work on this
6.1 Calculate relative formula mass given relative atomic masses			
6.2 Calculate the formulae of simple compounds from reacting masses and understand that these are empirical formulae			
HSW 11 Present information using scientific conventions and symbols			
6.3 Determine the empirical formula of a simple compound, such as magnesium oxide.			
6.4 Calculate the percentage composition by mass of a compound from its formula and the relative atomic masses of its constituent elements			
H 6.5 Use balanced equations to calculate masses of reactants and products			
HSW 10 Use qualitative and quantitative approaches when presenting scientific ideas and arguments, and recording observations			
6.6 Recall that the yield of a reaction is the mass of product obtained in the reaction			
6.7 Demonstrate an understanding that the actual yield of a reaction is usually less than the yield calculated using the chemical equation (theoretical yield)			
6.8 Calculate the percentage yield of a reaction from the actual yield and the theoretical yield			
6.9 Demonstrate an understanding of the reasons why reactions do not give the theoretical yield due to factors, including:			
a incomplete reactions			
b practical losses due to the preparation			
c competing, unwanted reactions			
HSW 12 Describe the benefits, drawbacks and risks of using new scientific and technological developments			
6.10 Demonstrate an understanding that many reactions produce waste products which:			
a are not commercially useful			
b can present economic, environmental and social problems for disposal			
H 6.11 demonstrate an understanding that chemists in industry work to find the economically most favourable reactions where:			
a the percentage yield is high			
b all the products of the reaction are commercially useful			
c the reaction occurs at a suitable speed			
HSW 13 Describe the social, economic and environmental effects of decisions about the uses of science and technology			